



STUDIES OF SOME UNI-UNIVALENT ION EXCHANGE REACTION SYSTEMS USING STRONGLY BASIC ANION EXCHANGE RESIN TULSION A-33

P.U.Singare*¹, R.S.Lokhande² and P.C.Vartak³

¹ Department of Chemistry, Bhavan's College, Andheri, Mumbai - 58

² Department of Chemistry, University of Mumbai, Vidyanagari, Santacruz, Mumbai-98

³ Department of Chemistry, VIVA College, M.B.Estate, Virar (West), Mumbai-4013030

E-mail:pravinsingare@vsnl.net

ABSTRACT

The thermodynamic study was carried to predict the selectivity behaviour of ion exchange resin Tulsion A-33 for inorganic anions like iodide and bromide. The equilibrium constant K values for the ion exchange reactions were calculated at different temperatures from which the enthalpy values were obtained. The equilibrium constant K calculated for Cl^- / I^- and $\text{Cl}^- / \text{Br}^-$ uni-univalent ion exchange reaction systems were observed to decrease with rise in temperature indicating the exothermic exchange reactions having enthalpy values -8.43 and -15.30 kJ/mol respectively.

Key words: Ion exchange equilibrium; Equilibrium constant; Enthalpy; Anion exchange; Tulsion A-33.

INTRODUCTION

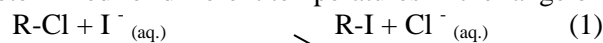
Ion exchange is one of the most common techniques that have been employed for many years in chemical process industries and effluent treatment plant. The ion exchange process is very effective at transferring the content of a large volume of industrial effluent into a small volume of solid. Extensive work was done by previous researchers to study the properties of the ion exchange resins, to generate thermodynamic data related to various uni-univalent and heterovalent ion exchange systems¹⁻⁷. Recently theories explaining ion exchange equilibrium between the resin phase and solution was also developed⁸. A number of researchers carried out equilibrium studies, extending over a wide range of composition of solution and resin phase⁹⁻³¹. Attempts were also made to study the equilibrium of cation exchange systems⁹⁻²³. However very little work was carried out to study the temperature effect on anion exchange systems^{12, 24-31} for computing the thermodynamic equilibrium constants. Therefore in the present investigation attempts were made to study the thermodynamics of uni-univalent anion exchange equilibrium, the results of which will be of considerable use in explaining the selectivity of ion exchanger for various univalent ions in solution.

EXPERIMENTAL

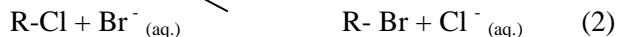
The ion exchange resin Tulsion A-33 as supplied by the manufacturer (Thermax Ltd., Pune) was a strongly basic anion exchange resin in hydroxide form of 16-50 mesh size. For present investigation, the resin grains of 30-40 mesh size were used. The conditioning of the resins in chloride form was done by usual methods using 10% potassium chloride solution²⁵⁻²⁹.

0.500g of ion exchange resins in Cl^- form was equilibrated with I^- ion solution of different concentrations at a constant temperature of 30.0 °C for 3 h. From the results of kinetics study reported earlier³²⁻⁴³, it was observed that this duration was adequate to attain the ion exchange equilibrium. After 3 h the different I^- ion solutions in equilibrium with ion exchange resins were analysed for their Cl^- and I^- ion concentration by potentiometric titration with standard 0.1N AgNO_3 solution. From the results the

equilibrium constant K for the reaction was determined at 30.0 °C. The equilibrium constants K for the above Cl⁻ / I⁻ system was determined for different temperatures in the range of 30.0 °C to 45.0 °C.



Similar study was also carried out for Cl⁻ / Br⁻ system in the same temperature range, to study the equilibrium constant K for the reaction



The chloride bromide and iodide ion solutions used in the entire experimental work, were prepared by dissolving their respective analytical grade potassium salts in distilled deionised water. In the present study, a semi-micro burette having an accuracy of 0.05 mL was used in the titrations and the titration readings were accurate to ± 0.05 mL. Considering the magnitude of the titer values, the average equilibrium constants reported in the experiment are accurate to ± 3 %.

RESULTS AND DISCUSSION

Therefore the equilibrium constants for the reactions (1 and 2) can be given by the expression

$$K = \frac{C_{\text{RX}} \cdot C_{\text{Cl}^{-}}}{(A - C_{\text{RX}}) \cdot C_{\text{X}^{-}}} \quad (3)$$

Here A is the ion exchange capacity of the resin, X⁻ represents I⁻ or Br⁻ ions.

For different concentrations of X⁻ ions in solution at a given temperature, K values were calculated from which average value of K for that set of experiment was calculated (Tables 1 and 2). Similar values of K were calculated for Cl⁻ / I⁻ and Cl⁻ / Br⁻ systems for different temperatures (Table 3). The enthalpy value for the ion exchange reactions 1 and 2 were calculated by plotting the graph of log K against 1 / T (Figure 1). Bonner and Pruett¹⁶ studied the temperature effect on uni-univalent exchanges involving some bivalent ions. In all bivalent exchanges, the equilibrium constant decreases with rise in temperature resulting in exothermic reactions. Similarly in the present investigation, for the uni-univalent exchange reactions the value of equilibrium constant decreases with rise in temperature giving negative enthalpy values (Table 3), indicating the exothermic ion exchange reactions. The low enthalpy and higher K values for Cl⁻ / I⁻ exchange as compared to that for Cl⁻ / Br⁻ exchange (Table 3), indicate that the resins in Cl⁻ form are having more affinity for larger ionic size I⁻ ions in solution as compared to that for Br⁻ ions also in the solution.

CONCLUSION

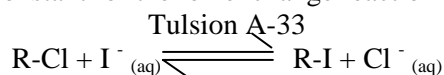
Various aspects of ion exchange technologies have been continuously studied to improve the efficiency and economy of ion exchangers in various technical applications. The selection of an appropriate ion exchange material for the specific industrial and research application is possible on the basis of information provided by the manufacturer. However, it is expected that the data obtained from the actual experimental trials will prove to be more helpful. The thermodynamic data obtained in the present experimental work will be useful to understand the selectivity behaviour of ion exchange resins for various ions in solution thereby helping in characterization of resins.

REFERENCES

1. N.N. Andreev and Y.I. Kuznetsov, *Russ. J. Phys. Chem.*, **64**, 1537 (1990).
2. A. Bhargava and C. Janardanan, *Indian J. Chem.*, **36A**, 624 (1997).
3. D. Muraviev, A. Gonzalo and M. Valiente, *Anal. Chem.*, **67**, 3028 (1995).
4. G.E. Boyd, F. Vaslow and S. Lindenbaum, *J. Phys. Chem.*, **71**, 2214 (1967).
5. J.F. Duncan, *Aus. J. Chem. Soc.*, **8**, 1 (1955).
6. G.E. Boyd, F. Vaslow and S. Lindenbaum, *J. Phys. Chem.*, **68**, 590 (1964).
7. A. Schwarz and G.E. Boyd, *J. Phys. Chem.*, **69**, 4268 (1965).
8. N.I. Gamayunov, *Russ. J. Phys. Chem.*, **64**, 1787 (1990).
9. G.E. Boyd and G.E. Myers, *J. Phys. Chem.*, **60**, 521 (1956).
10. O.D. Bonner, *J. Phys. Chem.*, **59**, 719 (1955).
11. O.D. Bonner, *J. Phys. Chem.*, **58**, 318 (1954).
12. S. Lindenbaum, C.F. Jumper and G.E. Boyd, *J. Phys. Chem.*, **63**, 1924 (1959).

13. K.A. Kraus and R.J. Raridon, *J. Phys. Chem.*, **63**, 1901 (1959).
14. O.D. Bonner and W.H. Payne, *J. Phys. Chem.*, **58**, 183 (1954).
15. W.J. Argersinger and A.W. Davidson, *J. Phys. Chem.*, **56**, 92 (1952).
16. O.D. Bonner and R.R. Pruett, *J. Phys. Chem.*, **63**, 1420 (1959).
17. O.D. Bonner and F.L. Livingston, *J. Phys. Chem.*, **60**, 530 (1956)
18. O.D. Bonner and L.L. Smith, *J. Phys. Chem.*, **61**, 326 (1957)
19. O.D. Bonner, C.F. Jumper and O.C. Rogers, *J. Phys. Chem.*, **62**, 250 (1958).
20. O.D. Bonner and L.L. Smith, *J. Phys. Chem.*, **61**, 1614 (1957).
21. J. Kielland, *J. Soc. Chem. Ind.*, **54**, 232 (1935).
22. A.P. Vanselow, *J. Am. Chem. Soc.*, **54**, 1307 (1932).
23. G.L. Gaines (Jr.) and H.C. Thomas, *J. Chem. Phys.*, **21**, 714 (1953).
24. K.A. Kraus, R.J. Raridon and D.L. Holcomb, *Chromatogr. J.*, **3**, 178 (1960).
25. R.S. Lokhande, P.U. Singare and A.B. Patil, *Russ. J. Phys. Chem. A*, **81**, 2059 (2007).
26. P.U. Singare, R.S. Lokhande and T.S. Prabhavalkar, *Bull. Chem. Soc. Ethiop.*, **22**, 415 (2008).
27. R.S. Lokhande and P. U. Singare, *J. Ind. Council Chem.*, **24**, 73 (2007).
28. R.S. Lokhande, P.U. Singare and A.R. Kolte, *Bull. Chem. Soc. Ethiop.*, **22**, 107 (2008).
29. K.G. Heumann and K. Baier, *Chromatographia*, **15**, 701 (1982).
30. O.D. Bonner, G. Dickel and H. Brummer, *Z. Physik. Chem.*, **25**, 81 (1960).
31. G.L. Starobinet, V.S. Soldatov and A.A. Krylova, *Russ. J. Phys. Chem.*, **41**, 194 (1967).
32. R.S. Lokhande, P.U. Singare and A.B. Patil, *Radiochim. Acta*, **95**, 111 (2007).
33. R. S. Lokhande and P. U. Singare, *Radiochim. Acta*, **95**, 173 (2007).
34. R.S. Lokhande, P.U. Singare and A.R. Kolte, *Radiochim. Acta*, **95**, 595 (2007).
35. R. S. Lokhande, P. U. Singare and M. H. Dole, *J. Nucl. Radiochem. Sci.*, **7**, 29 (2006).
36. R.S. Lokhande and P.U. Singare, *J. Porous Mater*, **15**, 253 (2008).
37. R.S. Lokhande, P.U. Singare and P. Karthikeyan, *Russ. J. Phys. Chem. A*, **81**, 1768 (2007).
38. R.S. Lokhande, P.U. Singare and M.H. Dole, *Radiochemistry*, **49**, 519 (2007).
39. P.U. Singare, R.S. Lokhande and A.B. Patil, *Radiochim. Acta*, **96**, 99 (2008).
40. R.S. Lokhande, P.U. Singare and S.R.D. Tiwari, *Radiochemistry*, **50**, 633 (2008).
41. R.S. Lokhande, P.U. Singare and T.S. Prabhavalkar, *Russ. J. Phys. Chem. A*, **82**, 1589(2008).
42. R.S. Lokhande, P.U. Singare and S.A. Parab, *Radiochemistry*, **50**, 642 (2008).
43. R.S. Lokhande, P.U. Singare and V.V. Patil, *Radiochemistry*, **50**, 638 (2008).

Table- 1: Equilibrium constant for the ion exchange reaction using ion exchange resin

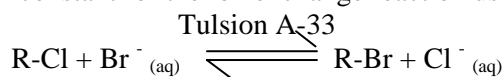


Amount of the ion exchange resin in Cl⁻ form = 0.500 g; Ion exchange capacity = 1.15 meq. / 0.500g;
Volume of I⁻ ion solution = 100.0 mL; Temperature = 35.0 °C

System	Initial conc. of I ⁻ ion (M)	Final conc. of I ⁻ ions (M) C _{I⁻}	Change in I ⁻ ion conc.	Conc. of Cl ⁻ ions exchanged (M) C _{Cl⁻}	Amount of I ⁻ ions exchanged on the resin meq./ 0.5 g C _{RI}	Equilibrium constant K
1	0.010	0.0003	0.0097	0.0096	0.485	15.29
2	0.020	0.0015	0.0185	0.0183	0.925	19.63
3	0.025	0.0023	0.0228	0.0226	1.138	31.52
4	0.030	0.0040	0.0260	0.0262	1.300	42.58

Average equilibrium constant (K) = 27.25

Table- 2: Equilibrium constant for the ion exchange reaction using ion exchange resin



Amount of the ion exchange resin in Cl⁻ form = 0.500 g; Ion exchange capacity = 1.15 meq. / 0.500g;
Volume of Br⁻ ion solution = 100.0 mL; Temperature = 35.0 °C

System	Initial conc. of Br ⁻ ions (M)	Final conc. of Br ⁻ ions (M) C _{Br⁻}	Change in Br ⁻ ion conc.	Conc. of Cl ⁻ ions exchanged (M) C _{Cl⁻}	Amount of Br ⁻ ions exchanged on the resin meq./ 0.5 g C _{RBr}	Equilibrium constant K
1	0.010	0.0008	0.0092	0.0091	0.460	5.03
2	0.020	0.0030	0.0170	0.0160	0.850	6.97
3	0.025	0.0045	0.0205	0.0204	1.025	9.78
4	0.030	0.0063	0.0238	0.0237	1.188	14.43

Average equilibrium constant (K) = 9.05

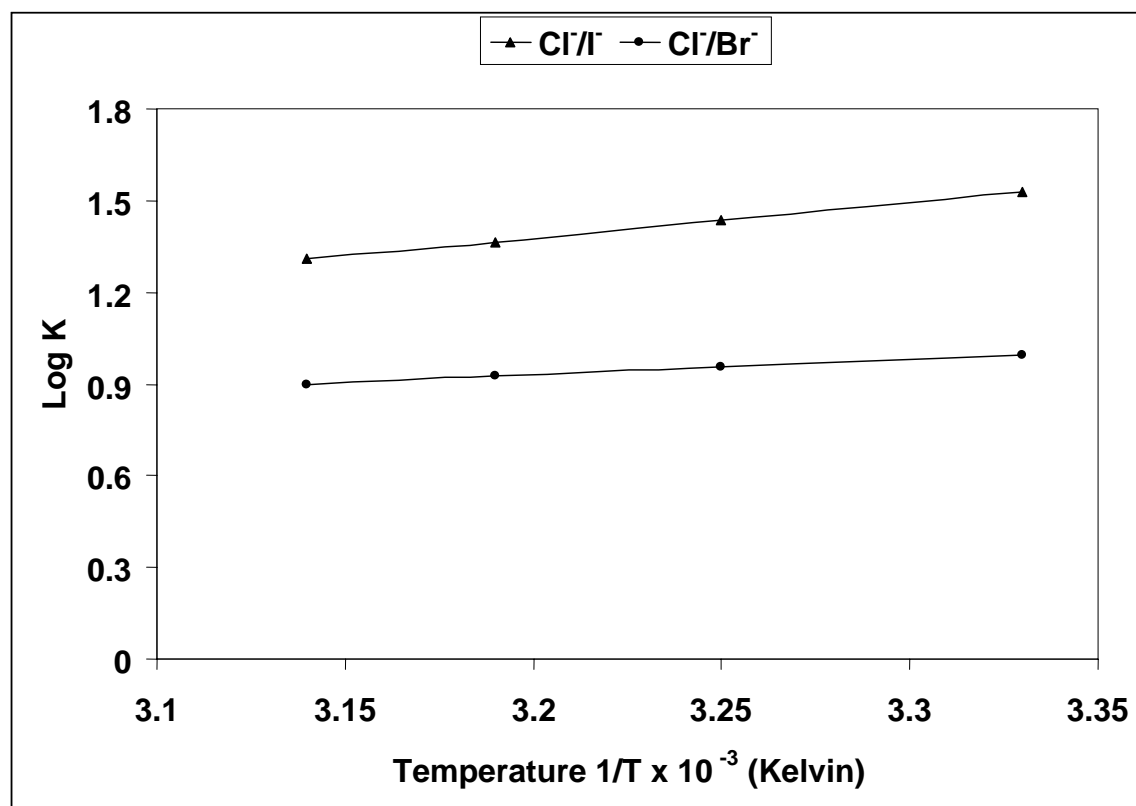


Fig.-1: Variation of Equilibrium Constant with Temperature for Uni-Univalent Ion Exchange Reactions using ion exchange resin Tulsion A-33

Amount of the ion exchange resin in Cl⁻ form = 0.500 g; Ion exchange capacity = 1.15 meq. / 0.500g;
Volume of I⁻ / Br⁻ ion solution = 100.0 mL; Temperature range = 30.0 °C-45.0 °C

Table -3: Effect of temperature on equilibrium constant for Uni-Univalent ion exchange reactions using ion exchange resin Tulsion A-33

Amount of the ion exchange resin in Cl⁻ form = 0.500 g; Ion exchange capacity = 1.15 meq. / 0.500g;
Volume of I⁻ / Br⁻ ion solution = 100.0 mL

Temp. °C	Equilibrium Constant (K) for the reactions	
	$R-Cl + I^{-} (aq.) \rightleftharpoons R-I + Cl^{-} (aq.)$	$R-Cl + Br^{-} (aq.) \rightleftharpoons R-Br + Cl^{-} (aq.)$
30.0	37.68	9.88
35.0	27.25	9.05
40.0	22.59	9.06
45.0	20.93	7.71
Enthalpy (kJ/mol)	-8.43	-15.30

(Received: 1 December 2008

Accepted: 4 January 2009

RJC-298)

“Without continual growth and progress, such words as improvement, achievement, and success have no meaning.”

-Benjamin Franklin